

**CHEMISTRY XI SCIENCE PART-I**

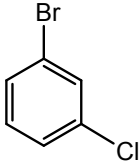
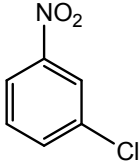
SL. No.	Page No.	Column	Line No.	Wrong	Corrected
1	4	L	38	polymesr	polymers
2	7	L	5	reaquired	required
3	10	R	15	$K=C+273.5$	$K=^{\circ}C+273.5$
4	10	R	17	below oC	below $0^{\circ}C$
5	11	R	29	$N 10n$	$N \times 10^n$
6	12	L	19	$6.65 \times 10^4$ and $8.95 \times 10^3$	$6.65 \times 10^4$ and $8.95 \times 10^3$
7	13	L	9	$4.01 \times 10^2$	$4.01 \times 10^2$
8	13	L	10	$8.256 \times 10^{-3}$	$8.256 \times 10^{-3}$
9	13	L	12	2.5 1.23	$2.5 \times 1.23$
10	14	R	29	duting	during
11	14	L	32	$2L=2 \times 1000\text{cm}^3$	$2L=2 \times 1000\text{cm}^3$
12	17	L	3	$1\text{amu}=1.66056 \times 10^{-24}\text{g}$	$1\text{amu}=1.66056 \times 10^{-24}\text{g}$
13	17	L	5	$1.6736 \times 10^{-24}\text{g}$	$1.6736 \times 10^{-24}\text{g}$
14	17	L	25	$14C \rightarrow 2 \times 10^{-10}$	$14C \rightarrow 2 \times 10^{-10}$
15	17	L	37	$2 \times 10^{-12}$	$2 \times 10^{-12}$
16	18	R	13	$6.022 \times 10^{23}$	$6.022 \times 10^{23}$
17	18	R	14	$6.022 \times 10^{23}$	$6.022 \times 10^{23}$
18	18	L	16	$1.992648 \times 10^{-23}$	$1.992648 \times 10^{-23}$
19	19	L	27	$(2 \times 12.01 + 6 \times 1.008 + 16.00)\text{g}$	$(2 \times 12.01 + 6 \times 1.008 + 16.00)\text{g}$
20	22	L	11	2 18	$2 \times 18$
21	22	L	14	2 18	$2 \times 18$
22	22	R	10	$17.86 \times 10^2$	$17.86 \times 10^2$
23	22	R	13	$4.96 \times 10^3$	$4.96 \times 10^3$

Sl. No.	Page No.	Column	Line No.	Incorrect	Corrected
24	22	R	19	$5.36 \times 10^3$	$5.36 \times 10^3$
25	22	R	20	$4.96 \times 10^3$	$4.96 \times 10^3$
26	22	R	24	$4.96 \times 10^3$	$4.96 \times 10^3$
27	22	R	26	$4.96 \times 10^3$	$4.96 \times 10^3$
28	22	R	27	$3.30 \times 10^3$	$3.30 \times 10^3$
29	22	R	28	$3.30 \times 10^3$	$3.30 \times 10^3$
30	22	R	32	$3.30 \times 10^3$	$3.30 \times 10^3$
31	23	L	1	$3.30 \times 10^3$	$3.30 \times 10^3$
32	23	L	2	$56.1 \times 10^3$	$56.1 \times 10^3$
33	23	R	32	$0.2\text{M } 1000\text{mL}=1.0\text{M } V_2$	$0.2\text{M} \times 1000\text{mL} = 1.0\text{M} \times V_2$
34	24	L	10	$0.4 \text{mol}^{-1}$	$0.4 \text{mol L}^{-1}$
35	24	R	8	3 58.5	$3 \times 58.5$
36	24	R	10	1000 1.25	$1000 \times 1.25$
37	24	R	12	1250-75.5	$1250-175.5$
38	37	R	28	duel	dual
39	39	R	18	$v = c/\lambda$	$v = c/\lambda$
40	43	R	3	$h\nu = hc/\lambda$	$E = hc/\lambda$
41	52	L	15	$10^{-4} \text{m}^{-2}\text{s}^{-1}$	$10^{-4} \text{m}^2\text{s}^{-1}$

42	79	L	12	n	v
43	111	L	23	$1D=3.33564 \times 10^{-30} \text{ C m}$	$1D=3.33564 \times 10^{-30} \text{ C m}$
44	111	R	11	$10^{-3} \text{ Cm}$ ( $1D=3.33564 \times 10^{-30} \text{ C m}$ )	$X 10^{-3} \text{ Cm}$ ( $1D=3.33564 \times 10^{-30} \text{ C m}$ )
45	111	R	21	$1.85 \times 3.33564 \times 10^{-30} \text{ C m}$	$1.85 \times 3.33564 \times 10^{-30} \text{ C m}$
46	115		4	$\text{SO}_2\text{O}_3$	$\text{SO}_2, \text{O}_3$

Sl. No.	Page No.	Column	Line No.	Wrong	Corrected
47	128	L	3	Antobpnding MOs $\sigma^*2s \sigma^*2p_z$ $\pi^*2p_x \pi^*2p_y$	Antibonding MOs $\sigma^*2s \sigma 2p_z$ $\pi^*2p_x \pi^*2p_y$
48	128	L	4	Bonding MOs $\sigma^*2s \sigma^*2p_z$ $\pi^*2p_x \pi^*2p_y$	Bonding MOs $\sigma 2s \sigma 2p_z \pi^*2p_x \pi^*2p_y$
49	129	L	4	$<(\pi^*2p_x = \pi^*2p_y) < \pi^*2p_z$	$<(\pi^*2p_x = \pi^*2p_y) < \sigma^*2p_z$
50	129	L	13	$<\sigma^*2p_z < (\pi^*2p_x = \pi^*2p_y) < \sigma^*2p_z$	$<\sigma 2p_z < (\pi^*2p_x = \pi^*2p_y) < \sigma^*2p_z$
51	130	L	40	$\text{C}_2: (\sigma 1s)^2 (\sigma^* 1s)^2 (\sigma^* 2s)^2 (\pi 2p_x^2 = \pi 2p_y^2)$	$\text{C}_2: (\sigma 1s)^2 (\sigma^* 1s)^2 (\sigma 2s)^2 (\sigma^* 2s)^2 (\pi 2p_x^2 = \pi 2p_y^2)$
52	130	R	28	$\text{O}_2 = (\sigma 1s)^2 (\sigma^* 1s)^2 (\sigma^* 2s)^2 (\sigma^* 2s)^2 (\sigma 2p_z)^2 (\pi 2p_x^2 = \pi 2p_y^2) (\pi^* 2p_x^1 = \pi^* 2p_y^1)$	$\text{O}_2 = (\sigma 1s)^2 (\sigma^* 1s)^2 (\sigma 2s)^2 (\sigma^* 2s)^2 (\sigma 2p_z)^2 (\pi 2p_x^2 = \pi 2p_y^2) (\pi^* 2p_x^1 = \pi^* 2p_y^1)$
53	130	R	30	$(\pi 2p_x^2 = \pi 2p_y^2) (\pi^* 2p_x^1 = \pi^* 2p_y^1)$	$(\pi 2p_x^2 = \pi 2p_y^2) (\pi^* 2p_x^1 = \pi^* 2p_y^1)$

CHEMISTRY XI SCIENCE PART II

Sl. No.	Page no	column	line no.	Wrong	Corrected
1	287	L	24	$2\text{NH}_2(\text{g})$	$2\text{NH}_3(\text{g})$
2		R	39	$\text{CO}(\text{g})+2\text{H}_2(\text{g})\rightarrow\text{CO}_3\text{OH}$	$\text{CO}(\text{g})+2\text{H}_2(\text{g})\rightarrow\text{CH}_3\text{OH}$
3	291	L	18	$\text{H}_2\text{O}+\text{H}_3\text{S}(\text{aq})$	$\text{H}_2\text{O}+\text{H}_2\text{S}(\text{aq})$
4		R	30	2 34	2 X 34
5	328	L	22	$\text{C}(\text{s})+\text{O}_2(\text{g})\rightarrow\text{CO}_2(\text{g})$	$\text{C}(\text{s})+\text{O}_2(\text{g})\rightarrow\text{CO}_2(\text{g})$
6	349	R	39	$\text{CO}(\text{g})+2\text{H}_2(\text{g})\rightarrow\text{CO}_3\text{OH}$	$\text{CO}(\text{g})+2\text{H}_2(\text{g})\rightarrow\text{CH}_3\text{OH}$
7	349	L	2	$\begin{array}{c} \text{H} \\ \{ \\ \text{CH}_3-\text{CH}_2-\text{C}=\text{O} \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{CH}_3-\text{CH}_2-\text{C}=\text{O} \end{array}$
8	351	R	15	$\text{HS}, \text{C}_2\text{H}_3\text{O}^-(\text{CH}_3\text{N}:$	$\text{HS}^-, \text{C}_2\text{H}_5\text{O}^-(\text{CH}_3)_3\text{N}:$
9		R	16	$\text{CH}_3, \text{C}_2\text{H}_3-\text{C}^+=\text{O}$	$\text{CH}_3^+, \text{C}_2\text{H}_5-\text{C}^+=\text{O}$
10		R		$\text{CH}_3, -\text{C}^+=\text{O}$	$\text{CH}_3^+, \text{C}_2\text{H}_5-\text{C}^+=\text{O}$
11	366	L		$2\text{NaOH}$	$\xrightarrow{\text{NaOH}}$
12	374	R	12	109.5	$109.5^0$
13	378	R	28	$\text{CH}_3-\text{C}=\text{C}-\text{H}$	$\text{CH}_3-\text{C}\equiv\text{C}-\text{H}$
14		R	34	$\xrightarrow{\text{Pt/Pd/Ni}}$	$\xrightarrow{\text{Zn, H}^+}$
15	382	L	9	$(\text{CH}_3\text{COO})_3$	$(\text{CH}_3\text{COO})_2\text{Mn}$
16		L	27	$\xrightarrow{\text{AnhyAlCl}_3/\text{HCl}}$	$\xrightarrow{\text{AnhyAlCl}_3/\text{HCl}}$
17		L	28	$-(\text{CH}_2)_2-\text{CH}_3+\text{CH}_3+\text{CH}_2+\text{CH}-$	$-(\text{CH}_2)_2-\text{CH}_3=\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_3$
18		R	17	$\text{CH}_4+\text{HO}_2$	$\text{CH}_4+\text{H}_2\text{O}$
19	391	R	4	$\xrightarrow{\text{KMnO}_4/\text{H}^+}$	$\xrightarrow{\text{KMnO}_4/\text{H}^+}$
20	391	R	5	$\xrightarrow{\text{KMnO}_4/\text{H}^+}$	$\xrightarrow{\text{KMnO}_4/\text{H}^+}$
21	395	L	13	$[\text{CH}_3=\text{CH}=\text{CH}_2]$	$[\text{CH}_3-\text{CH}=\text{CH}_2]$
22	414	R	17	$\xrightarrow{h\nu}\rightarrow\text{H}(\text{g})+\text{I}(\text{g})$	$\xrightarrow{h\nu}\rightarrow\text{OH}(\text{g})$
23	391	R	4	$\xrightarrow{h\nu}\rightarrow 2 \text{ I}(\text{g})$	$\xrightarrow{h\nu}\rightarrow 2 \text{ Cl}(\text{g})$
2	424				

## CHEMISTRY XII SCIENCE PART-I

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SL.N O	Pag e No	Col um n No.	Incorrect	Corrected
1	50	17	100 C	100°C
2	72	4	$E_{\text{Cu}^{2+}/\text{Cu}} = E_{\text{Cu}^{2+}/\text{Cu}}^0 - \frac{RT}{2F} \ln \frac{1}{[\text{Cu}^{2+}(\text{aq})]}$	$E_{\text{Cu}^{2+}/\text{Cu}} = E_{\text{Cu}^{2+}/\text{Cu}}^0 - \frac{RT}{2F} \ln \frac{1}{[\text{Cu}^{2+}(\text{aq})]}$
		9	$= E_{\text{Cu}^{2+}/\text{Cu}}^0 - E_{\text{Zn}^{2+}/\text{Zn}}^0 - \frac{RT}{2F} \ln \frac{1}{[\text{Cu}^{2+}(\text{aq})]} - \ln \frac{1}{[\text{Zn}^{2+}(\text{aq})]}$	$= E_{\text{Cu}^{2+}/\text{Cu}}^0 - E_{\text{Zn}^{2+}/\text{Zn}}^0 - \frac{RT}{2F} \left[ \ln \frac{1}{[\text{Cu}^{2+}(\text{aq})]} - \ln \frac{1}{[\text{Zn}^{2+}(\text{aq})]} \right]$
3	104	24	Alkeline medium	Alkaline medium
4	173	27	HEH angle()	HEH angle( <sup>0</sup> )
5	175	28	$(\text{NH}_4)_2\text{CO}_3 \rightleftharpoons 2\text{NH}_3 + \text{H}_2\text{O} + \text{CO}_2$	$(\text{NH}_4)_2\text{CO}_3 \rightarrow 2\text{NH}_3 + \text{H}_2\text{O} + \text{CO}_2$
6	176	5	200 10 <sup>5</sup> Pa	200 X 10 <sup>5</sup> Pa
	177		$\text{NH}_4\text{NO}_3 \xrightarrow{\text{Heat}} \text{N}_2\text{O} + 2\text{H}_2\text{O}$	$\text{NH}_4\text{NO}_3 \xrightarrow{\text{Heat}} \text{N}_2\text{O} + 2\text{H}_2\text{O}$
	178		$2 \text{Pb}(\text{NO}_3)_2 \xrightarrow{\text{Heat}} 4\text{NO}_2 + 2\text{PbO} + \text{O}_2$	$2 \text{Pb}(\text{NO}_3)_2 \xrightarrow{\text{Heat}} 4\text{NO}_2 + 2\text{PbO} + \text{O}_2$
7	180	27	$+\text{NO} \rightarrow [\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{+2} + \text{H}_2\text{O}$	$[\text{Fe}(\text{H}_2\text{O})_6]^{+2} + \text{NO} \rightarrow [\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{+2} + \text{H}_2\text{O}$
8	182	34	$3\text{CH}_3\text{COOH} + \text{PCl}_3 \rightarrow 3\text{CH}_3\text{CLCL} + \text{H}_3\text{PO}_3$	$3\text{CH}_3\text{COOH} + \text{PCl}_3 \rightarrow 3\text{CH}_3\text{COCl} + \text{H}_3\text{PO}_3$
9	227	14	E <sup>e</sup>	E <sup>0</sup>